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Thermochemistry

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162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write: $101.3 \text{ J} \cdot 0.18 \text{ L atm} \cdot 1 \text{ L atm}^{-1} = -18 \text{ J}$ 6.17 An expansion implies an increase in volume, therefore w must be -325 J (see the defining equation for pressure-volume work.)

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particles from an area of higher concentration to an area of lower concentration until the particles are evenly distributed. (2) iii) The ammonia solution gave off ammonia gas and

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184 Guided Reading and Study Workbook CHAPTER 17, Thermochemistry (continued) 6. In thermochemical calculations, is the direction of heat flow given from the point of view of the system, or of the surroundings? 7. What universal law states that energy can neither be created nor destroyed

SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)

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