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Worksheet

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present. In this
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given the mass
of $K_2Cr_2O_7$
in 1 mL of
solution, which
we can use to
calculate the
number of moles
of $K_2Cr_2O_7$...

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Answers Figure

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Answers

12.3. 45, 47,

48, 3/13/12:

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Complete all

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Stoichiometry

Example 1. How many molecules of SO_3 are

needed to react with 144 molecules of Fe_2O_3 given this balanced chemical

equation?. $\text{Fe}_2\text{O}_3 (\text{s}) + 3\text{SO}_3 (\text{g}) \rightarrow \text{Fe}_2(\text{SO}_4)_3$

4) 3. Solution.

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We use the balanced chemical equation to construct a conversion factor between Fe_2O_3 and SO_3 . The number of molecules of Fe_2O_3 goes on the bottom of our conversion factor so it

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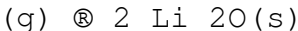
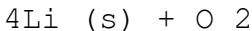
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Practice: 1. In

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a. what is the

product? Li_2O

(s) b. what are

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the reactants?

$\text{Li (s) O}_2 \text{ (g)}$

c. what does the
“(s)” after the

formula of

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